

## Lecture 05 Hydration

- a) Reaction of cement clinker
- b) Hydration modelling with GEMS

Barbara Lothenbach



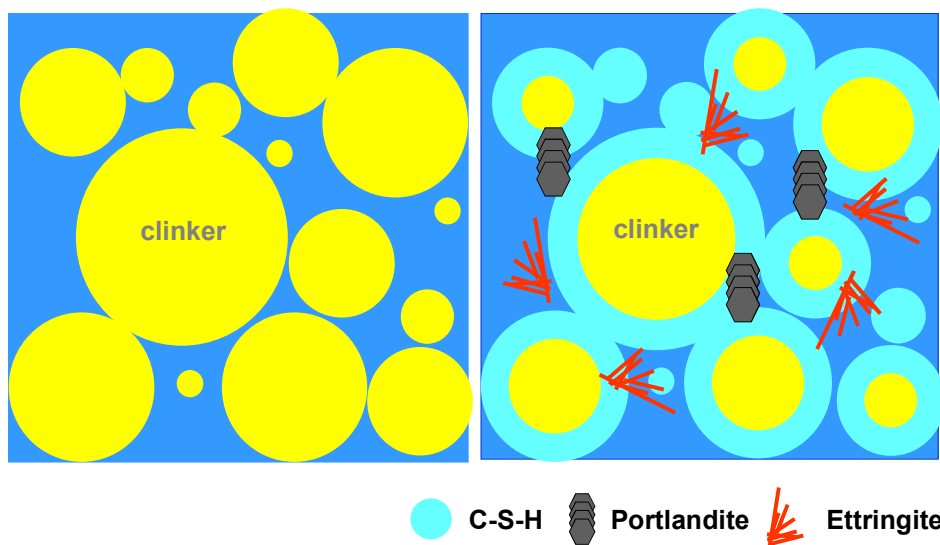
Software development/fitting  
tools/kinetic:

Dmitrii Kulik  
Dan Miron



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## Hydration



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# Modeling: Dissolution

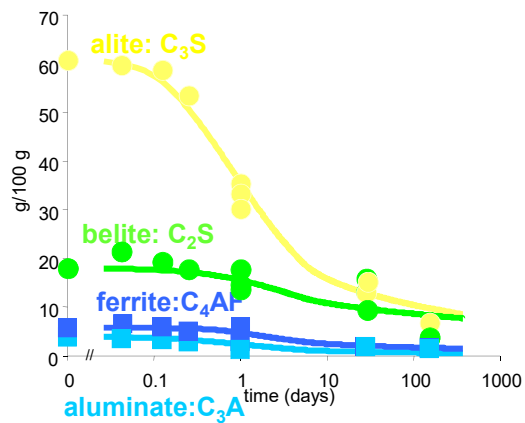
## Empirical Approach: Parrot and Killoh (1984)

$$R_t = \frac{K_1}{N_1} (1 - \alpha_t) (-\ln(1 - \alpha_t))^{(1-N_1)}$$

$$R_t = \frac{K_2 \times (1 - \alpha_t)^{2/3}}{1 - (1 - \alpha_t)^{1/3}}$$

$$R_t = K_3 \times (1 - \alpha_t)^{N_3}$$

All parameters ( $K_i$ ,  $N_i$ ) from  
Parrot and Killoh (1984)



Cement specific input: surface area, w/c, composition

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# Modeling: Dissolution

## Empirical Approach: Parrot and Killoh (1984)

$$R_t = \frac{K_1}{N_1} (1 - \alpha_t) (-\ln(1 - \alpha_t))^{(1-N_1)}$$

nucleation

$$R_t = \frac{K_2 \times (1 - \alpha_t)^{2/3}}{1 - (1 - \alpha_t)^{1/3}}$$

diffusion

$$R_t = K_3 \times (1 - \alpha_t)^{N_3}$$

shell

degree of hydration

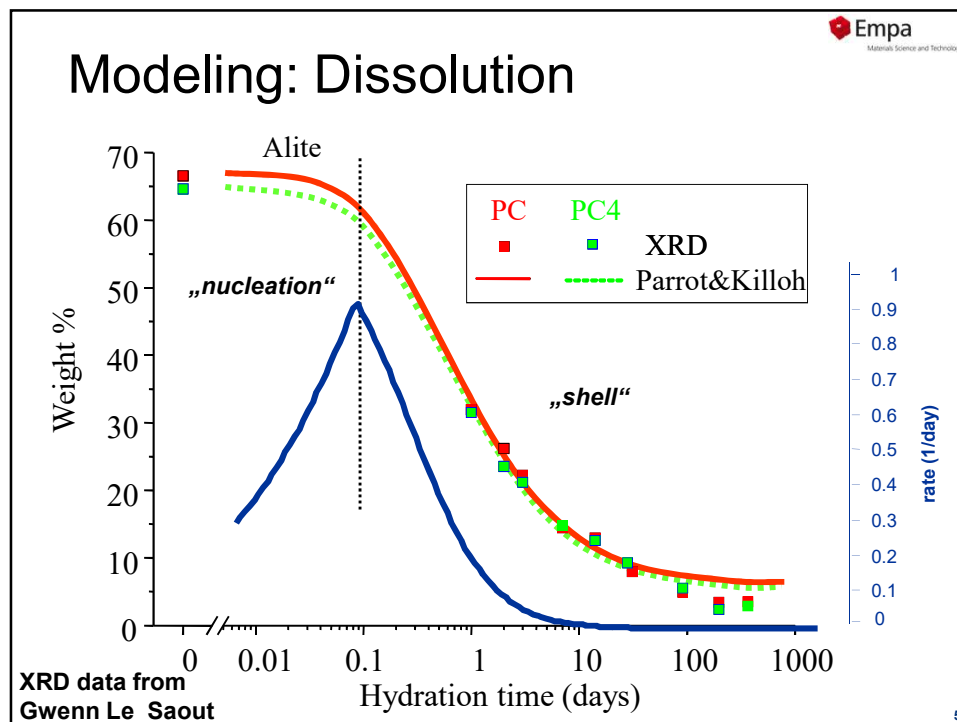
$$\alpha_t = \alpha_{t-1} + \Delta t \cdot R_{t-1} \cdot (1 + 3.333 \cdot (H \cdot w/c - \alpha_t))^4$$

for  $\alpha_t > H \cdot w/c$ .

	alite	belite	alum.	ferrite
$K_1$	1.5	0.5	1.0	0.37
$N_1$	0.7	1.0	0.85	0.7
$K_2$	0.05	0.006	0.04	0.015
$K_3$	1.1	0.2	1.0	0.4
$N_3$	3.3	5.0	3.2	3.7
H	1.33	1.33	1.33	1.33

Cement specific input:  
surface area, w/c

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Empa  
Materials Science and Technology

## Modeling: Dissolution

**Empirical Approach: Parrot and Killoh adapted**  
Lothenbach et al. (2008) 38, 848-860

*nucleation*

$$R_t = \frac{K_1}{N_1} (1 - \alpha_t) (-\ln(1 - \alpha_t))^{(1-N_1)}$$

*diffusion*

$$R_t = \frac{K_2 \times (1 - \alpha_t)^{2/3}}{1 - (1 - \alpha_t)^{1/3}}$$

*shell*

$$R_t = K_3 \times (1 - \alpha_t)^{N_3}$$

degree of hydration

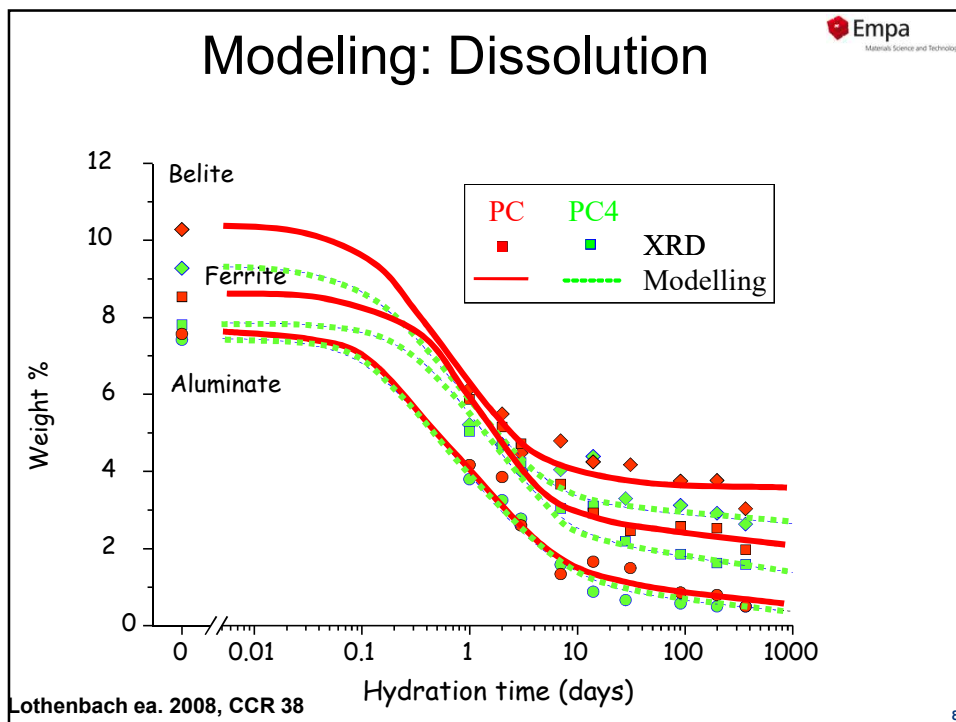
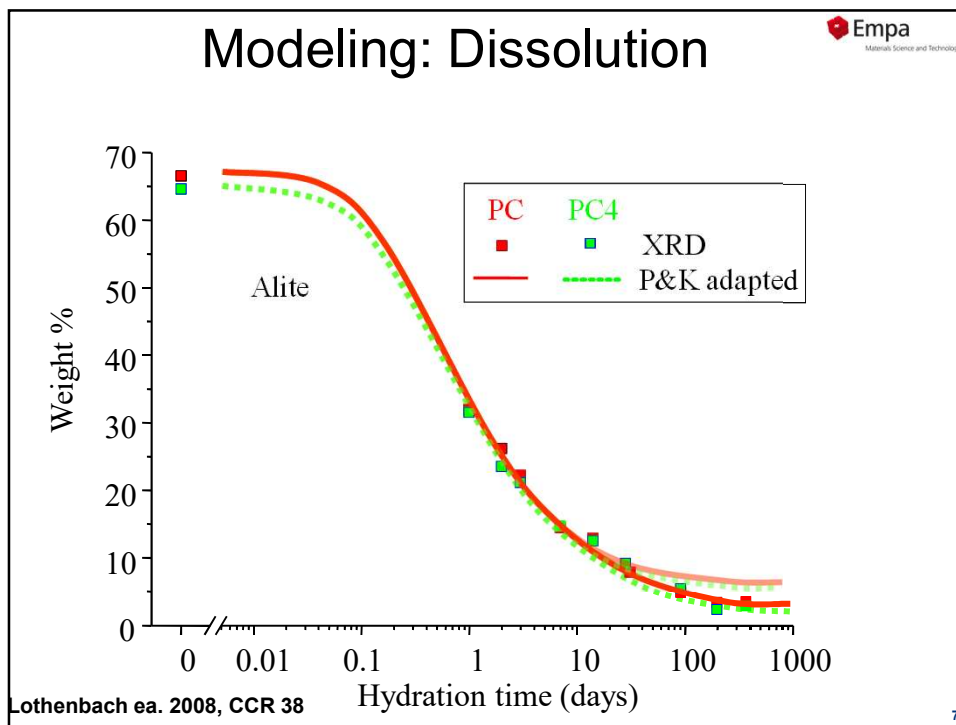
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	alite	belite	alum.	ferrite
$K_1$	1.5	0.5	1.0	0.37
$N_1$	0.7	1.0	0.85	0.7
$K_2$	0.05	<b>0.02</b>	0.04	0.015
$K_3$	1.1	<b>0.7</b>	1.0	0.4
$N_3$	3.3	5.0	3.2	3.7
H	<b>2.0</b>	<b>1.55</b>	<b>1.8</b>	<b>1.65</b>

**Cement specific input:  
surface area, w/c**

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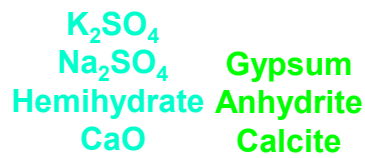
# Thermodyn. calculations Empa Materials Science and Technology

## Multi-component input

### I Slowly soluble clinkers



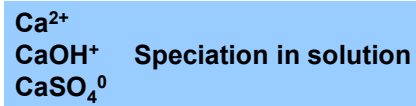
### II Soluble solids



### III Water



## Thermodynamic modeling GEMS-PSI



Portlandite



C-S-H



Ettringite



AFm

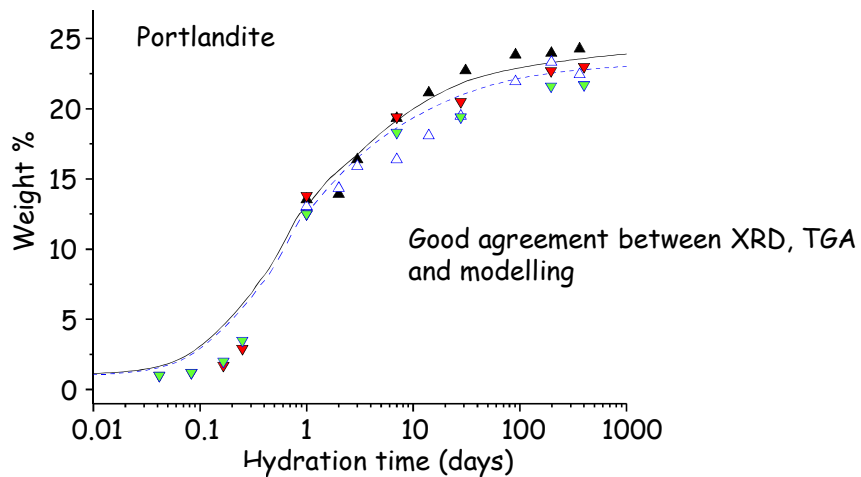


Hydrotalcites, ...

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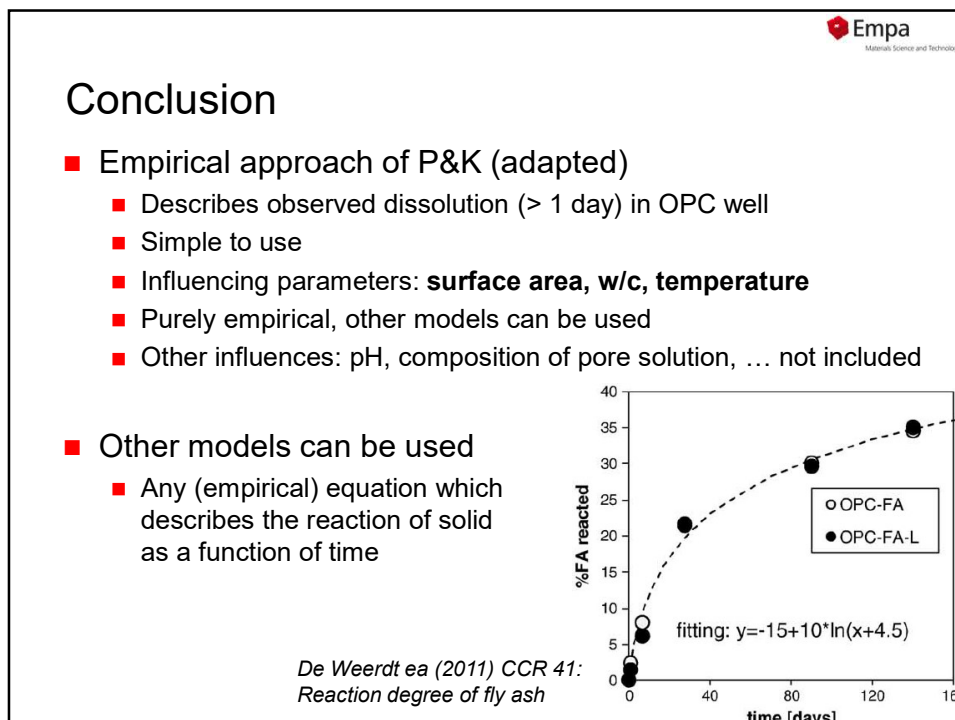
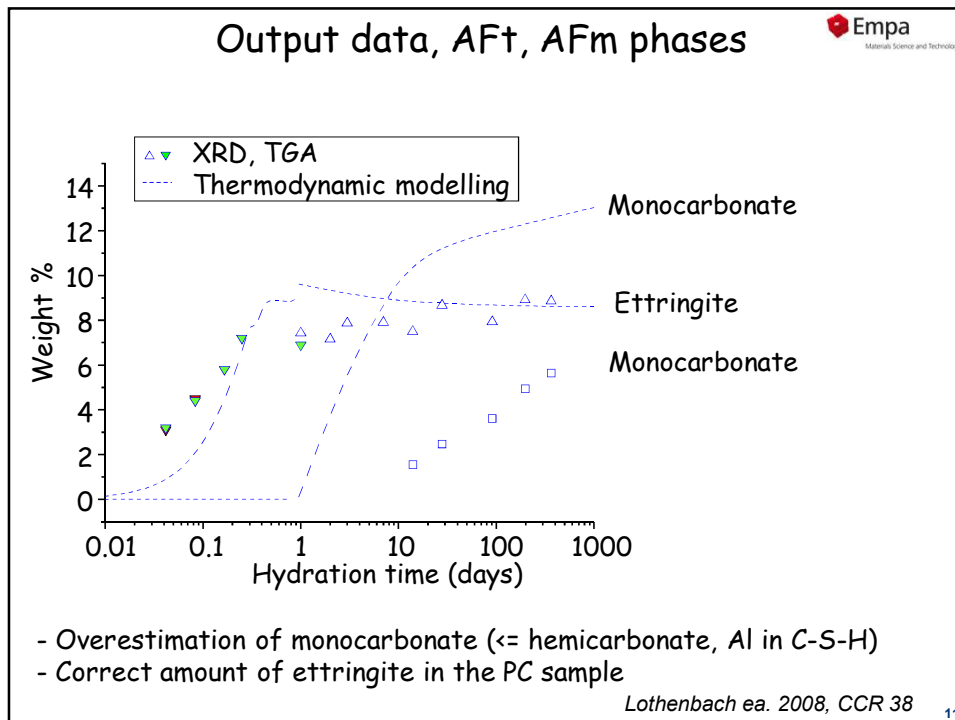
## Output data, Portlandite

PC	PC4
▲ ▼	△ ▽ XRD, TGA
—	- - - Thermodynamic modelling



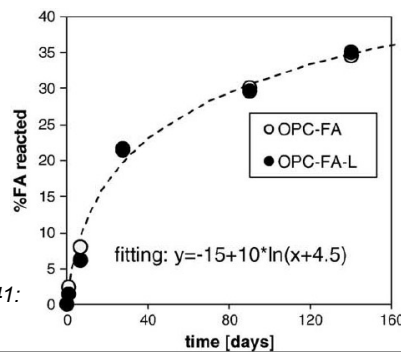
Lothenbach ea. 2008, CCR 38

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## Conclusion

- Empirical approach of P&K (adapted)
  - Describes observed dissolution (> 1 day) in OPC well
  - Simple to use
  - Influencing parameters: **surface area, w/c, temperature**
  - Purely empirical, other models can be used
  - Other influences: pH, composition of pore solution, ... not included
- Other models can be used
  - Any (empirical) equation which describes the reaction of solid as a function of time



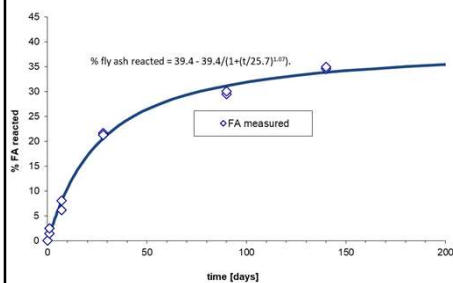
De Weerd et al (2011) CCR 41:  
Reaction degree of fly ash

## Alternative fits:

### ■ four (4PL)

$$\text{DoR} = D + (A - D) / (1 + (x/C)^B)$$

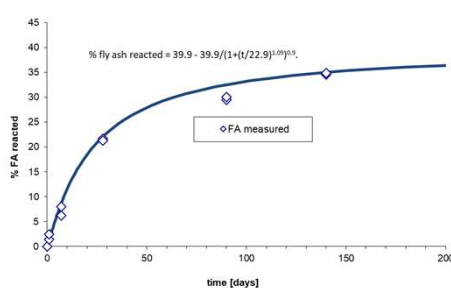
Symmetric



### five (5PL) parameter logistic fit

$$\text{DoR} = D + (A - D) / (1 + (x/C)^B)^G$$

Asymmetric, sometimes overfitting



A	minimum
D	maximum
C	inflection point
B	steepness of curve
G	asymmetry

## Modeling: Dissolution

$$R_{t,T} = \frac{K_1}{N_1} (1 - \alpha_t) (-\ln(1 - \alpha_t))^{(1-N_1)} \left( \frac{rh - 0.55}{0.45} \right)^4 \frac{\text{surface area}}{385} e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

$$R_{t,T} = \frac{K_2 \times (1 - \alpha_t)^{2/3}}{1 - (1 - \alpha_t)^{1/3}} \left( \frac{rh - 0.55}{0.45} \right)^4 e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

$$R_{t,T} = K_3 \times (1 - \alpha_t)^{N_3} \left( \frac{rh - 0.55}{0.45} \right)^4 e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

**Relative surface area factor**  
used for „nucleation and growth“ only  
(relative to Dalziel & Gutteridge, 1986)

**Influence of temperature**  
Arrhenius equation  
Ea values: Lothenbach et al., 2008

**Influence of the relative humidity**  
as proposed in Parrot and Killoh, 1984

**Hydration in closed systems: rh =1**

Cement specific input: surface area [m<sup>2</sup>/kg], w/c, composition

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## Modeling: Dissolution

$$R_{t,T} = \frac{K_1}{N_1} (1 - \alpha_t) (-\ln(1 - \alpha_t))^{(1-N_1)} \left( \frac{rh - 0.55}{0.45} \right)^4 \frac{\text{surface area}}{385} e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

$$R_{t,T} = \frac{K_2 \times (1 - \alpha_t)^{2/3}}{1 - (1 - \alpha_t)^{1/3}} \left( \frac{rh - 0.55}{0.45} \right)^4 e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

$$R_{t,T} = K_3 \times (1 - \alpha_t)^{N_3} \left( \frac{rh - 0.55}{0.45} \right)^4 e^{-\frac{E_a}{R} \left( \frac{1}{T} - \frac{1}{T_0} \right)}$$

degree of hydration of each clinker phase

- **initial**  $\alpha_t = \alpha_{t-1} + \Delta t \cdot R_{t-1}$
- **later, for  $\alpha_{t-1}(\text{total}) > H \cdot w/c$ ;**  
 $H \cdot w/c = \text{critical degree of hydration}$   
 $\alpha_t = \alpha_{t-1} + \Delta t \cdot R_{t-1} \cdot (1 + 3.333 \cdot (H \cdot w/c - \alpha_t))^4$

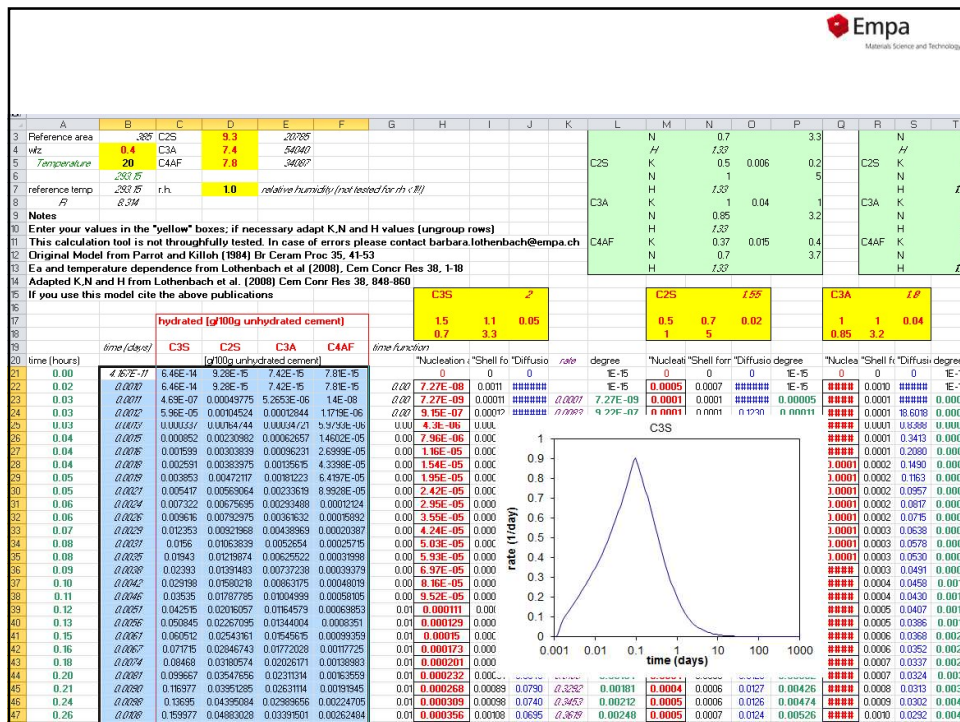
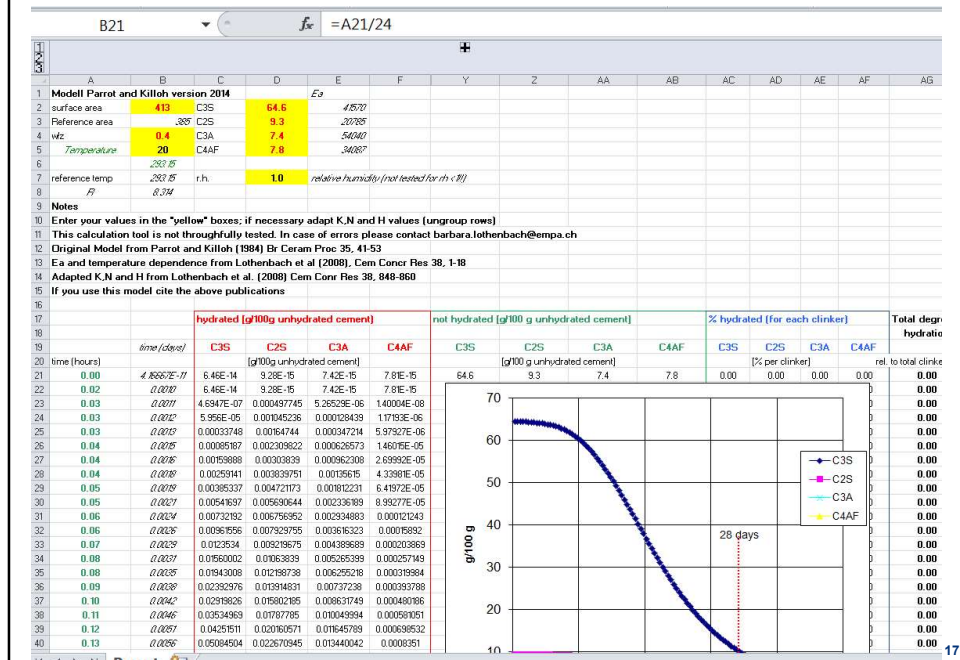
**Hydration is**  
**reduced with time**  
**at low w/c**

Cement specific input: surface area [m<sup>2</sup>/kg], w/c, composition

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# Parrot and Kiloh model as Excel file



# Hydration modelling EXCEL + GEMS

1. Copy folder **P\_K** into C:\Users\\*\*\*\Library\Gems3\projects
2. Open GEMS and project P\_K

**Parent file**

**Clinker phases incl. minor elements as predefined compositions**

Input Recipe of Single Thermodynamic System: Parrot\_K:G:parent:0:0:1:20:0

use for high C<sub>4</sub>AF, sulfate resistant cement

Property Selection

Compos (x<sub>a</sub>)

Al(OH)<sub>3</sub> CA2 Ferrite-HS Na2SO4  
Al2O3 CH4 Gypsum NaOH  
Al2Si2O5(OH)<sub>4</sub> O2 H2 O2  
Alite Ca(OH)<sub>2</sub> H2S PC  
Aluminate-c CaCO3 H2SO4 SO3  
Aluminate-o CaMg(CO3)<sub>2</sub> K2CO3 SiO2  
Aphthalite CaO K2O  
Belite CaSO4 K2SO4 Syngenite  
Belite CaSO4\_0.5H2O KOH  
C12A7 CaSiO3 Mg(OH)<sub>2</sub>  
C2S Fe2O3 MgSi2O5(OH)<sub>4</sub>  
C3A FeCO3 MgCO3  
C3S FeO MgO  
C4A3s FeOOH MgSO4  
C4AF FeS Na2CO3  
CA Ferrite Na2O

Recipe Input

Property	Name	Quantity	Units
1 x <sub>a</sub>	Alite	60	g
2 x <sub>a</sub>	Aluminate-c	5	g
3 x <sub>a</sub>	Aluminate-o	5	g
4 x <sub>a</sub>	Aqua	50	g
5 x <sub>a</sub>	Belite	10	g
6 x <sub>a</sub>	CaCO3	4	g
7 x <sub>a</sub>	Gypsum	4	g
8 x <sub>a</sub>	K2SO4	0.9	g
9 x <sub>a</sub>	MgO	1	g
10 x <sub>a</sub>	O2	1	g

Input quantities of Compos(tions) contributing to B<sub>-</sub> vector

additional phases aphthalite and syngenite sometimes present in PC clinkers

Print OK Cancel

All input in g  
(100 g cement)

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ergy Minimization - [Process :: Calculation finished OK (elapsed time: 14.138 s.)]

Parrot\_K:G:parent:0:0:1:20:0:PC\_ms:

Controls Sampling Results Config 04/09/2020, 13:37

CEM I

	iTm	iV	iP	iTC	iNv	iTau	ipXi	iNu	ipH	ipe
0	1000	0	1	20	0	0	0	0	0	0
1	1149	0	1	20	0	0	0	149		
2	1	0	0	0	0	0	0	1		
cTm	1149	0	1	20	0	0	0	149		

Reaction from Excel

Time (days)

Reacted C<sub>3</sub>S

g water per 100 g cement

allows blending with SCM (kinetics in modC[5])

```

$ General data
xa_[(Aqua)] = 40;
xa_[(O2)] = 0.1;
$ fraction of SCM added (0...1)
modC[20][6] = 0;
$ fraction of PC in the blend (1...0)
modC[21][6] = 1-modC[20][6];
$ Dissolution of clinker phases
xa_[(Alite)] = modC[J][1]*modC[0][6]*modC[21][6];
xa_[(Belite)] = modC[J][2]*modC[1][6]*modC[21][6];
xa_[(Aluminate-c)] = modC[J][3]*modC[2][6]*modC[21][6];
xa_[(Aluminate-o)] = modC[J][3]*modC[3][6]*modC[21][6];
xa_[(Ferrite)] = modC[J][4]*modC[4][6]*modC[21][6];
$ HgO same kinetics as C4AF
xa_[(MgO)] = modC[J][4]*modC[7][6]*modC[21][6];
$ Dissolution of SCM (example: microsilica)
xa_[(SiO2)] = modC[J][5]*100*modC[20][6];
$ Rapid soluble phases
xa_[(CaO)] = modC[5][6]*modC[21][6];
xa_[(Ca(OH)2)] = modC[6][6]*modC[21][6];
xa_[(CaSO4)] = modC[8][6]*modC[21][6];
xa_[(CaSO4_0.5H2O)] = modC[9][6]*modC[21][6];
xa_[(Gypsum)] = modC[10][6]*modC[21][6];
  
```

	modC[0]	modC[1]	modC[2]	modC[3]
0	4.16667e-011	1e-015	1e-015	
1	0.001	1e-015	1e-015	
2	0.001	7.26739e-005	0.00012633	1.7309
3	0.001	0.00012633	0.00012633	1.7309
4	0.001331	5.22414e-006	0.000177526	4.6794
5	0.0014641	1.31869e-005	0.000248903	8.4443
6	0.00161051	2.47504e-005	0.000327413	0.000
7	0.001771561	4.01147e-005	0.000413766	0.000
8	0.001948717	5.96498e-005	0.000508747	0.000
9	0.002143589	8.3854e-005	0.000613216	0.000
10	0.002357948	0.000113342	0.00072812	0.000
11	0.002593742	0.000148848	0.000854499	0.000
12	0.002853117	0.000191229	0.000993499	0.000
13	0.003138428	0.000241496	0.001146378	0.000
14	0.003452271	0.000300775	0.001314519	0.000
15	0.003797498	0.00037043	0.001499443	0.000
16	0.004177248	0.000451896	0.001702822	0.001
17	0.004594973	0.000547209	0.001926492	0.001
18	0.00505447	0.000658129	0.002172475	0.001

## Cement composition in modC[6]

### Reaction from Excel

Reaction degree

$C_3S$

$C_2S$

$C_3A$

$C_4AF$   
MgO

SCM

cement  
composition

	modC[1]	modC[2]	modC[3]	modC[4]	modC[5]	modC[6]	
0	1e-015	1e-015	1e-015	1e-015	2.18727e-009	64.6	C3S
1	1e-015	1e-015	1e-015	1e-015	0.000749251	9.3	C2S
2	7.26739e-009	5.36364e-005	7.09608e-007	1.79262e-009	0.00080471	3.7	C3A cubic
3	9.21989e-007	0.000112633	1.73099e-005	1.50055e-007	0.00086427	3.7	C3A orthorhombic
4	5.22414e-006	0.000177526	4.67943e-005	7.65592e-007	0.000928233	7.8	C4AF
5	1.31869e-005	0.000248903	8.44439e-005	1.86959e-006	0.000996923	0.5	CaO
6	2.47504e-005	0.000327413	0.000129691	3.457e-006	0.001070689	0.4	Ca(OH)2
7	4.01147e-005	0.000413766	0.00018277	5.5673e-006	0.001149905	0.9	MgO
8	5.96498e-005	0.000508747	0.000244236	8.21987e-006	0.001234972	3.1	gypsum
9	8.3854e-005	0.000613216	0.00031485	1.15144e-005	0.001326321	0	hemihydrate
10	0.000113342	0.00072912	0.000395537	1.55241e-005	0.001424414	0	anhydrite
11	0.000148848	0.000854499	0.000487375	2.03482e-005	0.001529747	1.2	K2SO4
12	0.000191229	0.000993499	0.000591602	2.61036e-005	0.001642852	0.2	Na2SO4
13	0.000241486	0.001146378	0.000709623	3.29256e-005	0.001764301	0	Aphthitalite
14	0.000300775	0.001314519	0.000843021	4.09711e-005	0.001894704	0	Syngenite
15	0.00037043	0.00149443	0.000993582	5.04209e-005	0.00203472	4.6	Calcite
16	0.000451986	0.001702822	0.001163308	6.14835e-005	0.002185052	0	Dolomite
17	0.000547209	0.001926492	0.001354447	7.43983e-005	0.002346457	0	Quartz
18	0.000658129	0.002172475	0.001569513	8.94407e-005	0.002519744	0	
19	0.000787075	0.00244299	0.001811326	0.000106927	0.002705782	0	

Data adapted from  
Lothenbach et al., 2008

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## Architecture of GEMS file

- Input always in g/100g
- The amount of clinker reacted copied from EXCEL file
- Inputs needed in GEMS:
  - w/c (as g water per 100 g cement)
  - Cement composition

C3S	64.6
C2S	9.3
C3A cubic	3.7
C3A orthorhombic	3.7
C4AF	7.8
CaO	0.5
Ca(OH)2	0.4
MgO	0.9
gypsum	3.1
hemihydrate	0
anhydrite	0
K2SO4	1.2
Na2SO4	0.2
Aphthitalite	0
Syngenite	0
Calcite	4.6
Dolomite	0
Quartz	0

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## Architecture of GEMS file

### ■ Process: sampling: definition of output

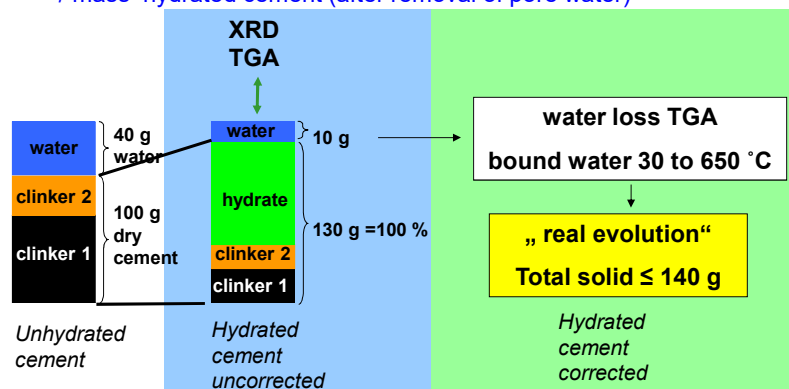
- X-axis: log (time)
- Mass in g/100 g of unhydrated cement
  - All data per 100 g unhydrated cement  
for comparison you will have to convert XRD/TGA data from g/100g hydrated paste to g/100g unhydrated cement
  - Amount of clinkers (with impurities)
  - Amount of hydrates (! Check single system file for the presence of additional solids and include them in the list)
  - Volumes in cm<sup>3</sup>/100g unhydrated cement (conversion needed for experimental data)
  - Amount of pore solution (including dissolved species)

**Check single system file for additional solids**

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## Different outputs possible

- Y-axis Mass in g/100 g of original cement  
 $\text{pH}M[\{\text{Portlandite}\}] \Rightarrow \text{correct measurements to 100 dry weight}$
- Y-axis Mass in g/100 g of hydrated cement  
 $\text{pH}M[\{\text{Portlandite}\}]/(1+\text{mod}C[1][5]-\text{pH}M[\{\text{aq\_gen}\}]/100)$   
 correction of the output  $/(1+0.4\text{-mass H}_2\text{O unreacted}/100)$   
 $\Rightarrow / \text{mass hydrated cement (after removal of pore water)}$



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## Different outputs possible

As the sample weight of the solid fraction of the sample is changing during hydration (see Figure 4.16 in Chapter 4 about XRD), the results need to be rescaled, to either paste or anhydrous, similarly as for the results of XRD-Rietveld analysis.

- Per 100 g paste:

$$\begin{aligned} \text{Ca(OH)}_{2,\text{paste}} &= \text{Ca(OH)}_{2,\text{measured}} / [(1 - \text{H}_2\text{O}_{\text{bound}})(1 + w/c)] \\ &= \frac{\text{Ca(OH)}_{2,\text{measured}}}{\text{weight at } 600^\circ\text{C}(1 + w/c)} \end{aligned} \quad (5.3)$$

- Per 100 g anhydrous:

$$\text{Ca(OH)}_{2,\text{dry}} = \text{Ca(OH)}_{2,\text{measured}} / (1 - \text{H}_2\text{O}_{\text{bound}}) = \frac{\text{Ca(OH)}_{2,\text{measured}}}{\text{weight at } 600^\circ\text{C}} \quad (5.4)$$

Lothenbach et al., Thermogravimetric analysis, in: Scrivener, Snellings, Lothenbach, A Practical Guide to Microstructural Analysis of Cementitious Materials, CRC Press, 2016

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tergy Minimization - [Process:: Calculation finished OK (elapsed time: 14:138 s)]

Controls | Sampling | Results | Config | 04/09/2020, 13:37

NeIt 9999 149 Next 0 I 0 J 148 Jp 148

pSTkey FK\_altern:G:parent:0:0:1:20:0: cTm 1148 cNV 0

cTau 0 cpXi 0 cXi 1 cNu 148

cpH 0 cpe 0 cEh 0 cT 293.15

```

xp[2] = lg(modC[2][0]);
$
$ Inert phases (dolomite, quartz)
yp[2][0] = (modC[16][6]*modC[17][6])*modC[21][6];
$
$ Clinker phases
yp[2][1] = modC[0][6]*modC[21][6]-ka_{{Alite}};
yp[2][2] = modC[1][6]*modC[21][6]-ka_{{Belite}};
yp[2][3] = modC[2][6]*modC[21][6]-ka_{{Aluminate-c}}+(modC[3][6]*modC[...
yp[2][4] = modC[4][6]*modC[21][6]-ka_{{Ferrite}};
yp[2][5] = modC[7][6]*modC[21][6]-ka_{{MgO}};
$
$ SCM (example: silica fume)
yp[2][6] = modC[20][6]*100-ka_{{SiO2}};
$ Hydrates
yp[2][7] = phM[{{CSHQ}}];
yp[2][8] = phM[{{Portlandite}}];
yp[2][9] = phM[{{ettringite}}]+phM[{{SO4_CO3_AFe}}]+phM[{{CO3_SO4_AFe}}];
yp[2][10] = phM[{{ettringite}}]+phM[{{SO4_CO3_AFe}}]+phM[{{CO3_SO4_AFe}}]+ph
yp[2][11] = phM[{{C3A}}];
yp[2][12] = phM[{{C4A}}];
yp[2][13] = phM[{{C3A}}];
yp[2][14] = phM[{{C4A}}];
yp[2][15] = phM[{{C3(A,F) SO.84H}}]+phM[{{C3F50.84H4.32}}]+phM[{{C3F51.34H3.32...
yp[2][16] = 0;
yp[2][17] = phM[{{Gypsum}}];

```

pX\_Nam lg t(d)

pLnam inert C3S C2S C3A C4A

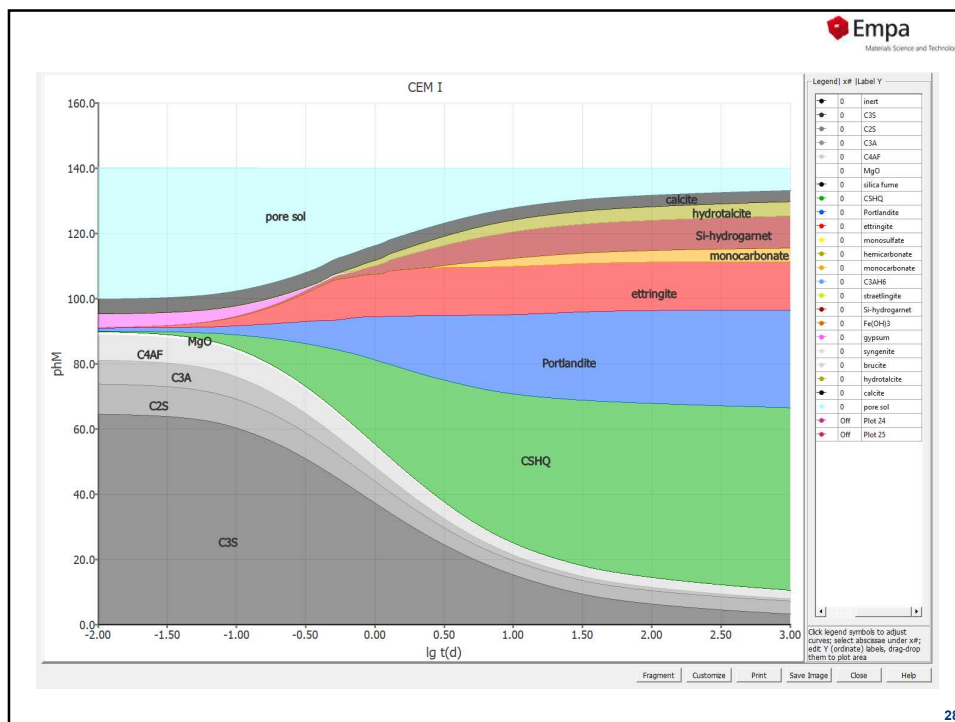
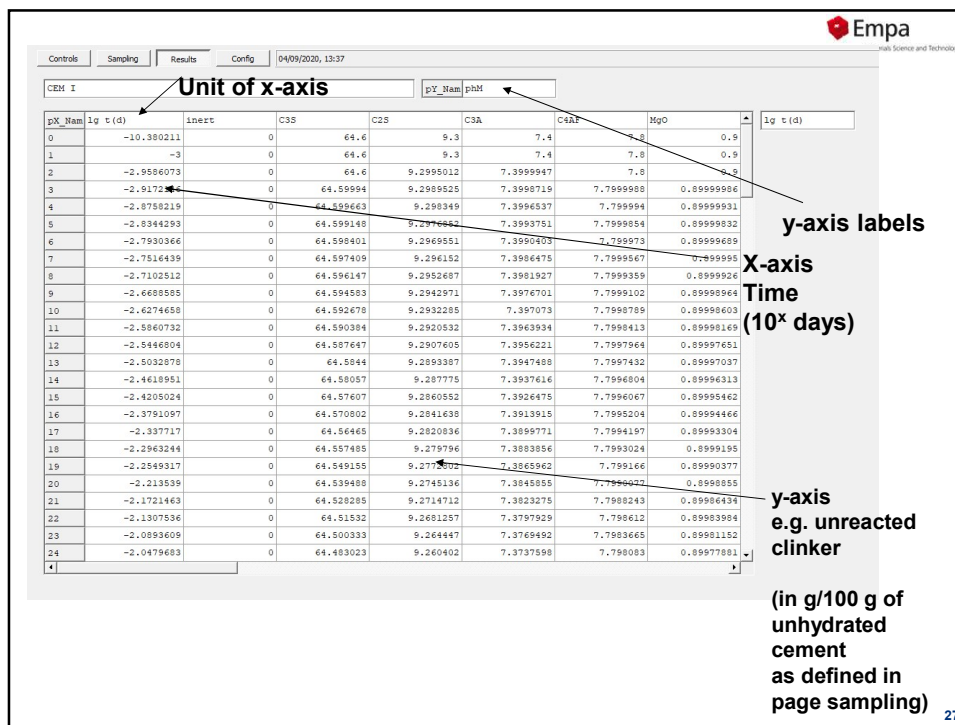
phM[{{CSHQ}}] -> mass of CSHQ in g per 100 g unhydrated cement  
! Spelling has to correspond exactly to single component!

Al/Fe-ss of ettringite deactivated due to stability issues

Option to deactivate siliceous hydrogarnet (deactivate also in parent file)

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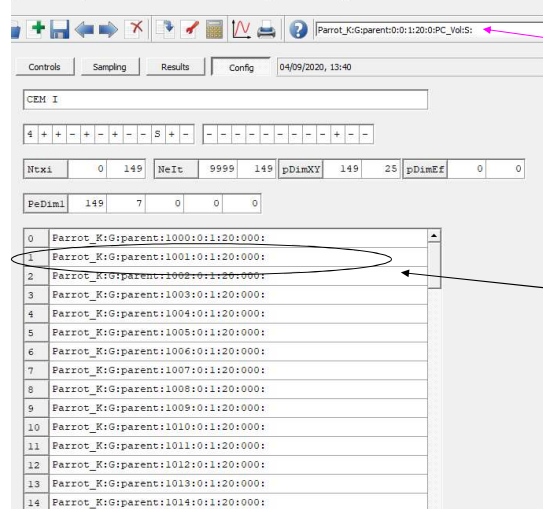




## Architecture of GEMS file

### ■ Process: Config: single system files

by Minimization - [Process: Definition of a Process Simulator (batch calculation)]



Name of the „parent file“  
in „Single-System Equilibria“

Names of the „kid files“  
produced:  
The results of each calculation  
can be checked

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## Architecture of GEMS file

- Process: Results: output
- X-axis: log (time)
  - Y-axis Mass in g/100 g of hydrated cement
  - Experimental data, same format as calculated data
    - Number of data points can be adapted by „Record:Remake“
- Data used to prepare graph, can be exported to Excel or other softwares by copy/paste

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### Number of experimental points

### Number of data categories

Selektor 3 (GEMS3) - Geochemical Equilibrium Model

Record Window Help

Process

 Parrot:\*,\*,\*,\*,\*,\*,\*,\*,\*,\*

	3	4	5	6	7	8	9
1	PC	0	0	1	20	0	Mass_corr
2	PC	0	0	1	20	0	mass

### Step 4 - Important data object dimensions

Spin boxes below define the dynamic memory configuration of the process

Dimensions of sampled and experimental data

149	nPS - Number of steps (1 to 9999 ) to be performed in this simulation
7	Number of 'modC' array columns (1 to 40, 0 - not used) to store nPS.
18	Number of columns in the 'yp' table (0 to 200) to keep the simulation of rows will be nPS.
1	Number of columns in the 'xp' table (0 to 4) to keep the simulation of rows will be nPS.
24	Number of rows in the xEp, yEp arrays for experimental data
9	Number of columns in the xEp, yEp arrays for experimental data

Optional data vectors (of length nPS) can be used for accumulating current data. They can be allocated using checkboxes below. The assignment operation automatically copied into data vector from the respective process iteration.

Allocation of optional data vectors

<input type="checkbox"/> CSD variant # ('vTm')	<input type="checkbox"/> Volume V, l ('vV')
<input type="checkbox"/> Temperature T ('vT')	<input type="checkbox"/> Constraints # ('vNV')
<input type="checkbox"/> Process extent pXi ('vpXi')	<input type="checkbox"/> Kinetic parameters ('vKin')

**Learn more**

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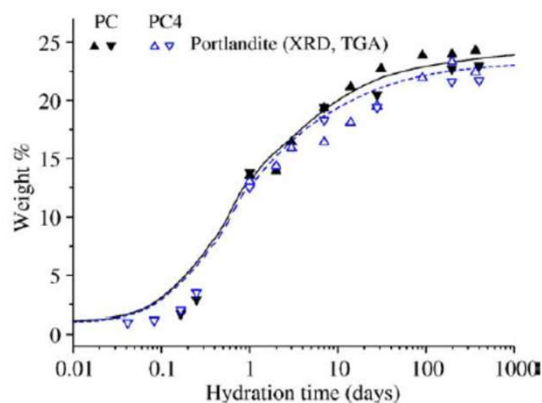


Fig. 11. Amounts of portlandite deduced by XRD/Rietveld analysis and thermal analysis as a function of hydration time. Lines refer to the results of thermodynamic modelling.

*Lothenbach et al., 2008*



## Comparison with experimental data

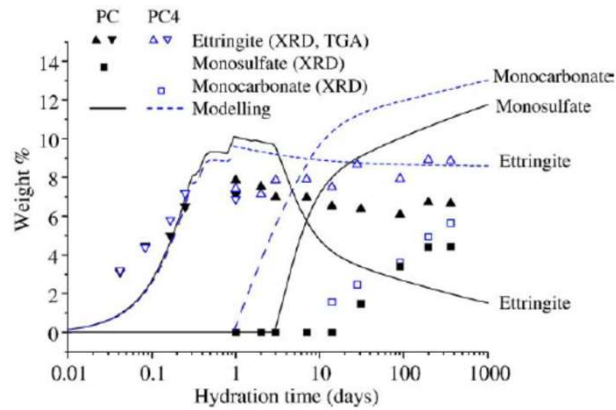


Fig. 12. Amounts of hydrated crystalline products deduced by XRD/ Rietveld analysis as a function of hydration time. Lines refer to the results of thermodynamic modelling.

Lothenbach et al., 2008

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## Comparison with experimental data

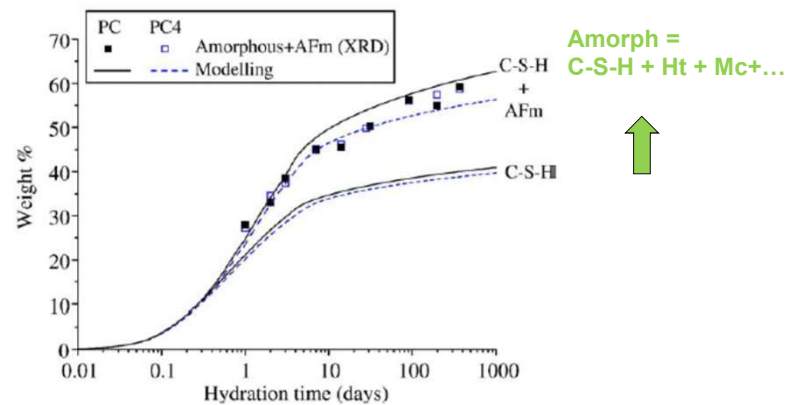


Fig. 13. Amounts of amorphous hydrated products deduced by XRD/ Rietveld analysis as a function of hydration time. Lines refer to the results of thermodynamic modelling.

Lothenbach et al., 2008

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## Sampling of aqueous concentrations

- Controls (= input) identical
- Page sampling and results adapted

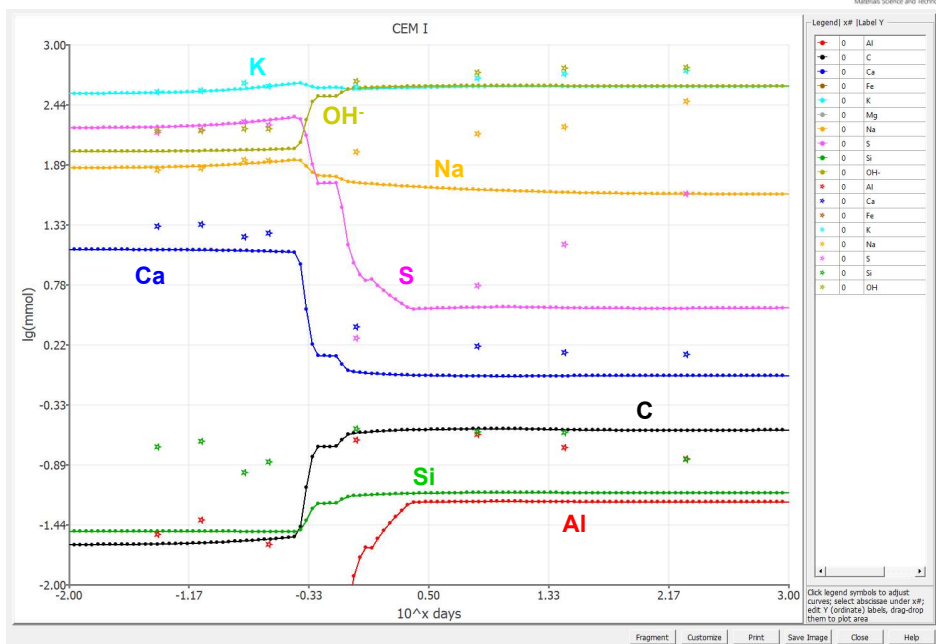
Controls	Sampling	Results	Config	04/09/2020, 13:37
NeIt	9999	149	Next	0
I	0	J	148	Jp
pSTkey	PK_altern:G:parent:0:0:1:20:0:			cTm
				1148
cTau	0	cpX1	0	cX1
				1
cNu				148
cpH	0	cpe	0	cEn
				0
cT				293.15

```

xp[J] =: lg(modC[J][0]);
$
$ Concentration (in mmol/l 1 solution)
$ re-calculation from molal (mol/kg H2O) -> molar (mol/l solution)
yp[J][0] =: lg(m_t[Al]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][1] =: lg(m_t[C]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][2] =: lg(m_t[Ca]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][3] =: lg(m_t[Fe]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][4] =: lg(m_t[K]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][5] =: lg(m_t[Mg]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][6] =: lg(m_t[Na]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][7] =: lg(m_t[S]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][8] =: lg(m_t[Si]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);
yp[J][9] =: lg(m_t[OH-]*x[{H2O}]*18.0153/phVol[{aq_gen}]*1000);

```

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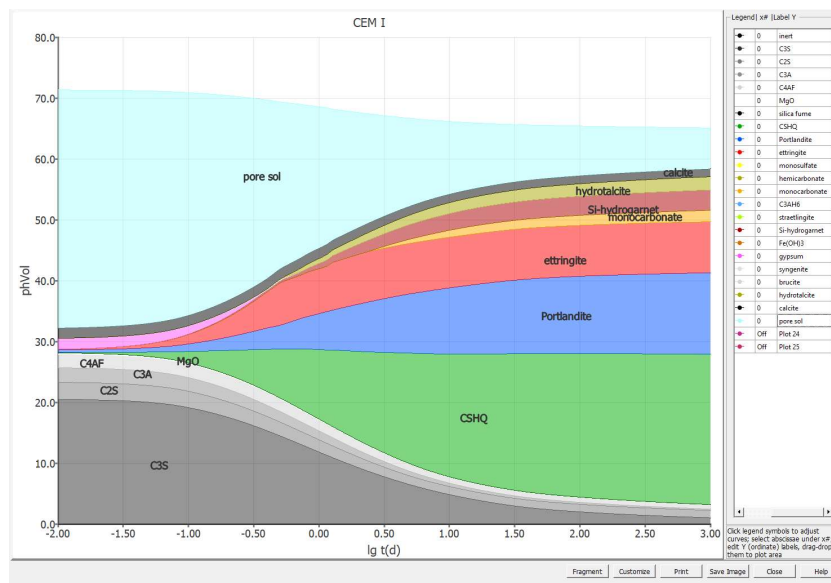
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## Excercise

- Calculate volumes of the hydrating cement
- Hint:
  - Duplicate process «mass»
  - Exchange pHM by pHVol
- Use density to calculate volume of unreacted clinker:
  - $yp[J][1] =: (modC[0][6]*modC[21][6]-xa_{[Alite]})/3.15;$
  - $yp[J][2] =: (modC[1][6]*modC[21][6]-xa_{[Belite]})/3.30;$
  - $yp[J][3] =: (modC[2][6]*modC[21][6]-xa_{[Aluminate-c]})/3.064+(modC[3][6]*modC[21][6]-xa_{[Aluminate-o]})/3.052;$
  - $yp[J][4] =: (modC[4][6]*modC[21][6]-xa_{[Ferrite]})/3.57;$
  - $yp[J][5] =: (modC[7][6]*modC[21][6]-xa_{[MgO]})/3.58;$

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### Volume in cm<sup>3</sup>/100 g unreacted cement



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